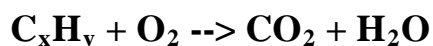




## TYPES OF REACTIONS

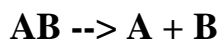
**Combustion** - A combustion reaction is when all substances in a compound are combined with oxygen, which then produces carbon dioxide and water. Combustion is commonly called burning. It is an exothermic reaction, which means heat is produced. An example of a combustion reaction is as follows:



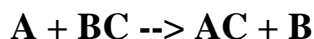
**Synthesis** - A synthesis reaction is when there is a combination of two or more substances and a new compound is made. An example of a synthesis reaction is as follows:



**Decomposition** - Decomposition is the opposite of synthesis. It is when a compound is broken down into simpler substances. An example of decomposition is as follows:



**Single Displacement** - In a single replacement reaction, one element of a compound is replaced with another element of a reacting compound. An example of a single replacement reaction is as follows:



**Double Displacement** - In a double replacement reaction, two elements of two different compounds are replaced (or exchanged) to yield two new compounds. An example of a double replacement reaction is as follows:

